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Tetraaquabis(6-carboxy-1*H*-benzimidazole-5-carboxylato-κN^3)nickel(II) dimethylformamide disolvate dihydrate

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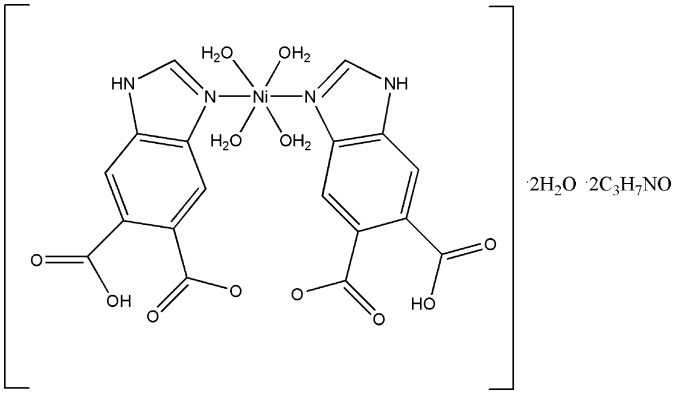
Key indicators: single-crystal X-ray study; T = 293 K; mean σ(C–C) = 0.004 Å; R factor = 0.037; wR factor = 0.137; data-to-parameter ratio = 12.6.

The title compound, $[Ni(C_9H_{45}N_2O_4)_2(H_2O)_4]\cdot 2C_3H_7NO\cdot 2H_2O$, has the Ni^{II} center coordinated by four water molecules and two N atoms from two 1*H*-benzimidazole-5,6-dicarboxylate ligands in an octahedral geometry. The molecule interacts with the solvent water and dimethylformamide molecules through N—H···O and O—H···O hydrogen bonds to form a three-dimensional supramolecular network. The metal atom lies on a center of inversion.

Related literature

For the crystal structures of 1*H*-benzimidazole-5,6-dicarboxylate complexes, see: Gao *et al.* (2008); Lo *et al.* (2007); Song *et al.* (2009).



Experimental

Crystal data

$[Ni(C_9H_5N_2O_4)_2(H_2O)_4]\cdot 2C_3H_7NO\cdot 2H_2O$
M_r = 723.30
Triclinic, $P\overline{1}$
a = 8.5327 (17) Å
b = 9.1387 (18) Å
c = 11.624 (2) Å
α = 100.80 (3)°
β = 103.03 (3)°
γ = 114.04 (3)°
V = 765.7 (3) Å^3
Z = 1
Mo $K\alpha$ radiation
μ = 0.72 mm^{-1}
T = 293 K
0.27 × 0.18 × 0.17 mm

Data collection

Rigaku/MSC Mercury CCD diffractometer
Absorption correction: multi-scan (*REQAB*; Jacobson, 1998)
T_{min} = 0.830, T_{max} = 0.888
6116 measured reflections
2737 independent reflections
2613 reflections with $I > 2\sigma(I)$
R_{int} = 0.020

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.037
$wR(F^2)$ = 0.137
S = 1.20
2737 reflections
217 parameters
9 restraints
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.61 e Å^{-3}
$\Delta\rho_{min}$ = −0.32 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O1W—H2W···O1^{i}	0.84	1.86	2.693 (3)	173
O1W—H1W···O3^{ii}	0.84	2.00	2.801 (3)	160
O4—H4A···O5^{iii}	0.82	1.78	2.585 (3)	167
O2W—H4W···O2^{iv}	0.84	1.79	2.624 (3)	176
O2W—H3W···O1W^{v}	0.84	1.92	2.741 (2)	166
O3W—H5W···O1W^{v}	0.84	2.06	2.810 (3)	148
O3W—H6W···O1^{vi}	0.84	1.81	2.634 (3)	169
N2—H2···O5^{vii}	0.86	1.98	2.779 (3)	155

Symmetry codes: (i) $-x+1, -y+2, -z+1$; (ii) $x, y, z+1$; (iii) $-x+1, -y+2, -z$; (iv) $x-1, y-1, z$; (v) $-x+1, -y+1, -z+1$; (vi) $-x+2, -y+2, -z+1$; (vii) $-x, -y+1, -z$.

Data collection: *RAPID-AUTO* (Rigaku, 1998); cell refinement: *RAPID-AUTO*; data reduction: *CrystalStructure* (Rigaku/MSC, 2002); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEPII* (Johnson, 1976); software used to prepare material for publication: *SHELXL9*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NG2646).

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supplementary materials

Acta Cryst. (2009). E**65**, m1258 [**doi:10.1107/S1600536809038069**]

Tetraaquabis(6-carboxy-1*H*-benzimidazole-5-carboxylato-κN^3)nickel(II) dimethylformamide disolvate dihydrate

H. Wang, W.-D. Song, S.-J. Li, P.-W. Qin and S.-W. Hu

Comment

From the structuralpoint of view, 1*H*-benzimidazole-5,6-dicarboxylic acid possesses two nitrogen atoms of imidazole ring and four oxygen atoms of carboxylate groups, and might be used as versatile linker in constructing coordination polymers with abundant hydrogen bonds. And several coordination polymers fomed by this ligand have been reported recently:Pentaaqua(1*H*-benzimidazole-5,6-dicarboxylato-kN^3)copper(II) pentahydrate(Gao *et al.*,2008), Bis(1*H*-benzimidazole-5,6-dicarboxylato)bis[tetraaquadicobalt(II)] pentahydrate(Lo *et al.*,2007), Pentaaqua(1*H*-benzimidazole-5,6-dicarboxylato-kN^3) cobalt(II)pentahydrate(Song *et al.*,2009).In the present paper, we synthesized a novel coordination complex $[Ni(C_9H_4N_2O_4)_2(H_2O)_4].2H_2O.2C_3H_7NO$.

As shown in Figure 1, the Ni^{II} atom exhibits an octahedral coordination sphere, defined by two N atoms from two different 1*H*-benzimidazole-5,6-dicarboxylate ligands, and four water molecules. The equatorial plane is defined by O2w, O3w, $O2w^i$ and $O3w^i$ atoms, while N1 and $N1^i$ occupy the axial position (symmetry codes: i = 1 - *x*, 1 - *y*, 1 - *z*). Inter/intramolecular O—H···O and N—H···O hydrogen bonds between the carboxylate O atoms of 1*H*-benzimidazole-5,6-dicarboxylate and the coordinated water molecule lead to the structure more stable(Fig 2).The hydrogen bonds are in the normal range(Table 1).

Experimental

A C_3H_7NO solution (20 mL)containing $Ni(NO_3)_2$(0.1 mmol)and 1*H*-benzimidazole-5,6-dicarboxylic acid(0.2 mmol) was stirred for a few minutes in air,and left to stand at room temperature for about four weeks, then the green crystals were obtained.

Refinement

Carbon and nitrogen bound H atoms were placed at calculated positions and were treated as riding on the parent C or N atoms with C—H = 0.93 Å, N—H = 0.86 Å, and with U_{iso}(H) = 1.2 U_{eq}(C, N). The water H-atoms were located in a difference map, and were refined with a distance restraint of O—H = 0.84 Å; their U_{iso} values were refined.

Figures

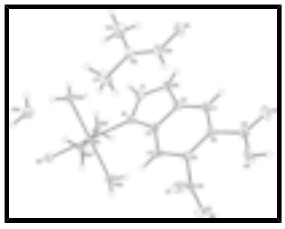

Fig. 1. The structure of the title compound, showing the atomic numbering scheme with 30% probability displacement ellipsoids. [Symmetry codes: (i) 1 - *x*, 1 - *y*, 1 - *z*.]

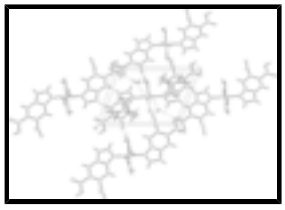

Fig. 2. A view of the three-dimensional network constructed by O—H···O and N—H···O hydrogen bonding interactions.

Tetraaquabis(6-carboxy-1*H*-benzimidazole-5-carboxylato- κ*N*^{3})nickel(II) dimethylformamide disolvate dihydrate

Crystal data

$[Ni(C_9H_5N_2O_4)_2(H_2O)_4]\cdot 2C_3H_7NO\cdot 2H_2O$	$Z = 1$
$M_r = 723.30$	$F_{000} = 378$
Triclinic, $P\overline{1}$	$D_x = 1.569$ Mg m^{-3}
Hall symbol: -P 1	Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
$a = 8.5327$ (17) Å	Cell parameters from 3600 reflections
$b = 9.1387$ (18) Å	$\theta = 1.4$–28º
$c = 11.624$ (2) Å	$\mu = 0.72$ mm^{-1}
$\alpha = 100.80$ (3)º	$T = 293$ K
$\beta = 103.03$ (3)º	Block, green
$\gamma = 114.04$ (3)º	0.27 × 0.18 × 0.17 mm
$V = 765.7$ (3) Å^3	

Data collection

Rigaku/MSC Mercury CCD diffractometer	2737 independent reflections
Radiation source: fine-focus sealed tube	2613 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{int} = 0.020$
$T = 293$ K	$\theta_{max} = 25.2$º
ω scans	$\theta_{min} = 3.3$º
Absorption correction: multi-scan (REQAB; Jacobson, 1998)	$h = -10\rightarrow10$
$T_{min} = 0.830$, $T_{max} = 0.888$	$k = -10\rightarrow9$
6116 measured reflections	$l = -13\rightarrow13$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.037$	H-atom parameters constrained
$wR(F^2) = 0.137$	$w = 1/[\sigma^2(F_o^2) + (0.1051P)^2 + 0.01P]$ where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.20$	$(\Delta/\sigma)_{max} < 0.001$
2737 reflections	$\Delta\rho_{max} = 0.61$ e Å^{-3}
217 parameters	$\Delta\rho_{min} = -0.32$ e Å^{-3}
9 restraints	Extinction correction: none
Primary atom site location: structure-invariant direct methods	

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
C1	0.2895 (3)	0.4030 (3)	0.2261 (2)	0.0301 (5)
H1	0.2030	0.3062	0.2339	0.036*
N1	0.4387 (2)	0.5153 (2)	0.31852 (17)	0.0272 (4)
Ni1	0.5000	0.5000	0.5000	0.02200 (19)
C2	0.4271 (3)	0.5938 (3)	0.1432 (2)	0.0272 (5)
N2	0.2739 (2)	0.4415 (2)	0.11945 (17)	0.0324 (4)
H2	0.1861	0.3832	0.0499	0.039*
C3	0.5296 (3)	0.6391 (3)	0.2686 (2)	0.0252 (4)
O3W	0.75996 (16)	0.54366 (16)	0.50333 (12)	0.0345 (4)
C4	0.6923 (3)	0.7892 (3)	0.3224 (2)	0.0271 (5)
H4	0.7612	0.8210	0.4056	0.033*
C5	0.7498 (3)	0.8909 (3)	0.2494 (2)	0.0248 (4)
O1	0.9319 (2)	1.1775 (2)	0.37262 (18)	0.0456 (5)
C6	0.6429 (3)	0.8423 (3)	0.1223 (2)	0.0280 (5)
O2	1.0690 (2)	1.0376 (2)	0.3063 (2)	0.0522 (5)
C7	0.4811 (3)	0.6925 (3)	0.0690 (2)	0.0307 (5)
H7	0.4113	0.6596	−0.0141	0.037*

C8	0.6944 (3)	0.9451 (3)	0.0384 (2)	0.0327 (5)
C9	0.9321 (3)	1.0500 (3)	0.31293 (19)	0.0261 (5)
O2W	0.41483 (16)	0.24688 (15)	0.42925 (12)	0.0295 (4)
H6W	0.8553	0.6369	0.5347	0.044*
H5W	0.7568	0.4874	0.4361	0.044*
H3W	0.4755	0.2330	0.3845	0.044*
H4W	0.3028	0.1830	0.3922	0.044*
O3	0.6082 (3)	0.8962 (3)	−0.07127 (17)	0.0591 (6)
O4	0.8343 (3)	1.0927 (3)	0.09199 (18)	0.0633 (7)
H4A	0.8481	1.1446	0.0415	0.095*
O1W	0.3424 (2)	0.7464 (3)	0.69171 (17)	0.0487 (5)
H1W	0.3996	0.7866	0.7687	0.073*
H2W	0.2637	0.7784	0.6716	0.073*
O5	0.0718 (2)	0.7407 (2)	0.05904 (16)	0.0462 (5)
C12	−0.0319 (4)	0.5721 (4)	0.2255 (3)	0.0570 (8)
H12A	−0.1264	0.6015	0.2316	0.086*
H12B	−0.0161	0.5128	0.2835	0.086*
H12C	−0.0653	0.5013	0.1426	0.086*
N3	0.1376 (3)	0.7248 (3)	0.25427 (18)	0.0352 (5)
C11	0.2581 (4)	0.8066 (4)	0.3847 (2)	0.0522 (7)
H11A	0.3625	0.9074	0.3914	0.078*
H11B	0.2971	0.7308	0.4129	0.078*
H11C	0.1933	0.8351	0.4351	0.078*
C10	0.1722 (3)	0.7970 (3)	0.1704 (2)	0.0356 (5)
H10	0.2794	0.8978	0.1949	0.043*

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0264 (11)	0.0250 (11)	0.0280 (11)	0.0029 (9)	0.0068 (9)	0.0092 (9)
N1	0.0272 (9)	0.0227 (9)	0.0259 (9)	0.0057 (7)	0.0082 (7)	0.0098 (7)
Ni1	0.0196 (3)	0.0191 (3)	0.0217 (3)	0.00424 (18)	0.00625 (17)	0.00652 (17)
C2	0.0216 (10)	0.0230 (10)	0.0262 (11)	0.0024 (8)	0.0062 (8)	0.0055 (9)
N2	0.0256 (9)	0.0274 (10)	0.0254 (9)	−0.0008 (8)	0.0021 (7)	0.0071 (8)
C3	0.0242 (10)	0.0248 (10)	0.0264 (11)	0.0093 (8)	0.0112 (8)	0.0099 (9)
O3W	0.0231 (8)	0.0295 (8)	0.0380 (9)	0.0049 (6)	0.0105 (6)	0.0003 (7)
C4	0.0231 (10)	0.0266 (11)	0.0223 (10)	0.0059 (8)	0.0037 (8)	0.0060 (8)
C5	0.0223 (10)	0.0237 (11)	0.0245 (10)	0.0075 (8)	0.0074 (8)	0.0078 (8)
O1	0.0282 (9)	0.0305 (9)	0.0559 (11)	0.0043 (7)	0.0102 (7)	−0.0071 (8)
C6	0.0252 (10)	0.0278 (11)	0.0258 (11)	0.0073 (9)	0.0082 (8)	0.0098 (9)
O2	0.0233 (9)	0.0324 (9)	0.0833 (15)	0.0073 (7)	0.0126 (8)	0.0002 (9)
C7	0.0283 (11)	0.0312 (11)	0.0217 (10)	0.0069 (9)	0.0037 (8)	0.0072 (9)
C8	0.0295 (11)	0.0328 (12)	0.0266 (12)	0.0063 (9)	0.0078 (9)	0.0113 (10)
C9	0.0220 (10)	0.0249 (11)	0.0249 (11)	0.0059 (9)	0.0050 (8)	0.0095 (9)
O2W	0.0250 (7)	0.0233 (7)	0.0316 (8)	0.0058 (6)	0.0076 (6)	0.0054 (6)
O3	0.0561 (12)	0.0527 (12)	0.0269 (9)	−0.0082 (9)	0.0002 (8)	0.0191 (9)
O4	0.0599 (12)	0.0464 (11)	0.0327 (10)	−0.0161 (9)	−0.0030 (9)	0.0224 (9)
O1W	0.0396 (10)	0.0590 (12)	0.0390 (10)	0.0258 (9)	0.0087 (8)	−0.0036 (9)

O5	0.0409 (10)	0.0468 (10)	0.0272 (9)	0.0019 (8)	0.0025 (7)	0.0156 (8)
C12	0.0560 (17)	0.0591 (18)	0.0607 (19)	0.0195 (14)	0.0312 (15)	0.0326 (15)
N3	0.0421 (11)	0.0367 (11)	0.0260 (10)	0.0184 (9)	0.0101 (9)	0.0103 (9)
C11	0.078 (2)	0.0531 (17)	0.0285 (13)	0.0379 (16)	0.0077 (13)	0.0140 (12)
C10	0.0363 (12)	0.0301 (12)	0.0302 (12)	0.0083 (10)	0.0065 (10)	0.0104 (10)

Geometric parameters (Å, °)

C1—N1	1.317 (3)	C6—C7	1.386 (3)
C1—N2	1.344 (3)	C6—C8	1.492 (3)
C1—H1	0.9300	O2—C9	1.236 (3)
N1—C3	1.397 (3)	C7—H7	0.9300
N1—Ni1	2.1014 (18)	C8—O3	1.209 (3)
Ni1—O2W	2.0501 (14)	C8—O4	1.293 (3)
Ni1—O2W^{i}	2.0501 (14)	O2W—H3W	0.8401
Ni1—O3W^{i}	2.0773 (13)	O2W—H4W	0.8400
Ni1—O3W	2.0773 (13)	O4—H4A	0.8200
Ni1—N1^{i}	2.1014 (18)	O1W—H1W	0.8408
C2—C7	1.379 (3)	O1W—H2W	0.8405
C2—N2	1.390 (3)	O5—C10	1.252 (3)
C2—C3	1.401 (3)	C12—N3	1.454 (4)
N2—H2	0.8600	C12—H12A	0.9600
C3—C4	1.391 (3)	C12—H12B	0.9600
O3W—H6W	0.8402	C12—H12C	0.9600
O3W—H5W	0.8394	N3—C10	1.298 (3)
C4—C5	1.391 (3)	N3—C11	1.470 (3)
C4—H4	0.9300	C11—H11A	0.9600
C5—C6	1.424 (3)	C11—H11B	0.9600
C5—C9	1.522 (3)	C11—H11C	0.9600
O1—C9	1.240 (3)	C10—H10	0.9300
N1—C1—N2	113.71 (18)	C4—C5—C9	115.96 (18)
N1—C1—H1	123.1	C6—C5—C9	123.63 (19)
N2—C1—H1	123.1	C7—C6—C5	120.7 (2)
C1—N1—C3	104.99 (18)	C7—C6—C8	115.74 (19)
C1—N1—Ni1	123.63 (15)	C5—C6—C8	123.59 (19)
C3—N1—Ni1	131.30 (15)	C2—C7—C6	117.85 (19)
O2W—Ni1—O2W^{i}	180.0	C2—C7—H7	121.1
O2W—Ni1—O3W^{i}	91.85 (6)	C6—C7—H7	121.1
O2W^{i}—Ni1—O3W^{i}	88.15 (6)	O3—C8—O4	122.2 (2)
O2W—Ni1—O3W	88.15 (6)	O3—C8—C6	122.5 (2)
O2W^{i}—Ni1—O3W	91.85 (6)	O4—C8—C6	115.3 (2)
O3W^{i}—Ni1—O3W	180.0	O2—C9—O1	125.6 (2)
O2W—Ni1—N1^{i}	90.06 (7)	O2—C9—C5	116.7 (2)
O2W^{i}—Ni1—N1^{i}	89.94 (7)	O1—C9—C5	117.58 (18)
O3W^{i}—Ni1—N1^{i}	90.14 (7)	Ni1—O2W—H3W	109.0
O3W—Ni1—N1^{i}	89.86 (7)	Ni1—O2W—H4W	117.7

O2W—Ni1—N1	89.94 (7)	H3W—O2W—H4W	110.9
O2W^{i}—Ni1—N1	90.06 (7)	C8—O4—H4A	109.5
O3W^{i}—Ni1—N1	89.86 (7)	H1W—O1W—H2W	111.4
O3W—Ni1—N1	90.14 (7)	N3—C12—H12A	109.5
N1^{i}—Ni1—N1	180.0	N3—C12—H12B	109.5
C7—C2—N2	131.9 (2)	H12A—C12—H12B	109.5
C7—C2—C3	122.56 (19)	N3—C12—H12C	109.5
N2—C2—C3	105.48 (19)	H12A—C12—H12C	109.5
C1—N2—C2	106.79 (18)	H12B—C12—H12C	109.5
C1—N2—H2	126.6	C10—N3—C12	121.1 (2)
C2—N2—H2	126.6	C10—N3—C11	120.7 (2)
C4—C3—N1	131.2 (2)	C12—N3—C11	117.8 (2)
C4—C3—C2	119.76 (19)	N3—C11—H11A	109.5
N1—C3—C2	109.02 (19)	N3—C11—H11B	109.5
Ni1—O3W—H6W	126.3	H11A—C11—H11B	109.5
Ni1—O3W—H5W	111.1	N3—C11—H11C	109.5
H6W—O3W—H5W	111.6	H11A—C11—H11C	109.5
C5—C4—C3	118.77 (19)	H11B—C11—H11C	109.5
C5—C4—H4	120.6	O5—C10—N3	124.8 (2)
C3—C4—H4	120.6	O5—C10—H10	117.6
C4—C5—C6	120.39 (19)	N3—C10—H10	117.6

Symmetry codes: (i) −*x*+1, −*y*+1, −*z*+1.

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O1W—H2W···O1^{ii}	0.84	1.86	2.693 (3)	173
O1W—H1W···O3^{iii}	0.84	2.00	2.801 (3)	160
O4—H4A···O5^{iv}	0.82	1.78	2.585 (3)	167
O2W—H4W···O2^{v}	0.84	1.79	2.624 (3)	176
O2W—H3W···O1W^{i}	0.84	1.92	2.741 (2)	166
O3W—H5W···O1W^{i}	0.84	2.06	2.810 (3)	148
O3W—H6W···O1^{vi}	0.84	1.81	2.634 (3)	169
N2—H2···O5^{vii}	0.86	1.98	2.779 (3)	155

Symmetry codes: (ii) −*x*+1, −*y*+2, −*z*+1; (iii) *x*, *y*, *z*+1; (iv) −*x*+1, −*y*+2, −*z*; (v) *x*−1, *y*−1, *z*; (i) −*x*+1, −*y*+1, −*z*+1; (vi) −*x*+2, −*y*+2, −*z*+1; (vii) −*x*, −*y*+1, −*z*.

Fig. 1

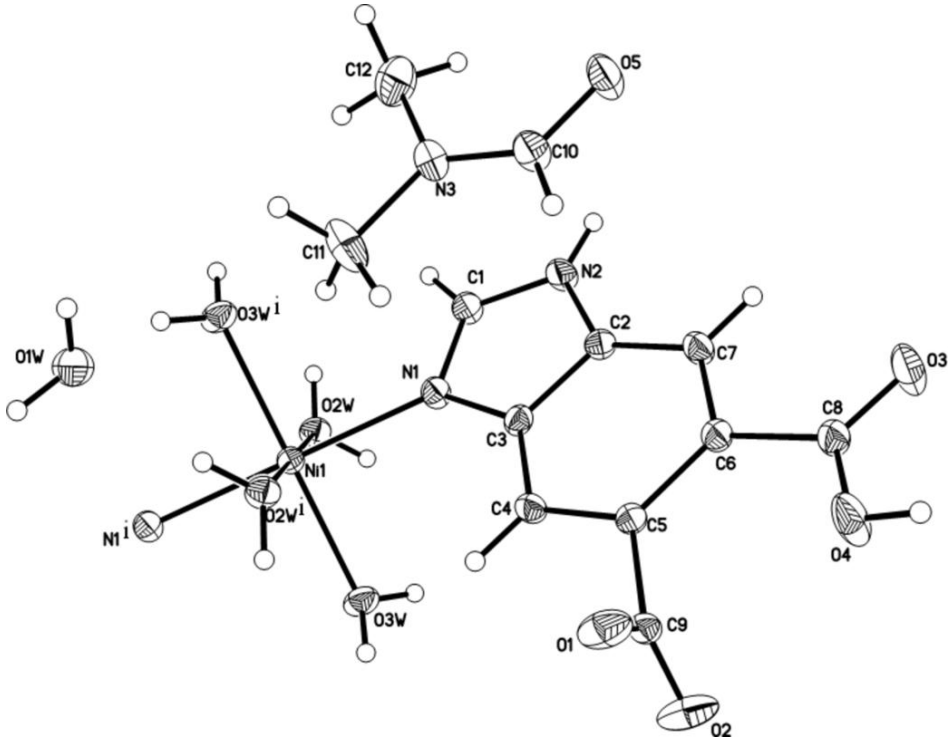
C12
O5
C10
N3
C11
C1
N2
O3W i
C2
O1W
C7
O3
N1
O2W
C8
C3
Ni1
C6
C4
C5
N1 i
O2W i
O4
O3W
O1
C9
O2

Fig. 2